



IDENTIFYING DATA

Chemistry: Chemistry 2

Subject	Chemistry: Chemistry 2			
Code	V11G200V01204			
Study programme	(*)Grao en Química			
Descriptors	ECTS Credits	Choose	Year	Quadmester
	6	Basic education	1st	2nd
Teaching language	Spanish Galician			
Department	Physical Chemistry Inorganic Chemistry Organic Chemistry			
Coordinator	Pérez Juste, Ignacio			
Lecturers	Castro Fojo, Jesús Antonio Hervés Beloso, Juan Pablo Pérez Juste, Ignacio Silva López, Carlos			
E-mail	uviqpij@uvigo.es			
Web	http://faic.uvigo.es			
General description	Chemistry II pretends to provide to students the basis for the understanding of disciplines more specific, that will give in future courses.			

Competencies

Code	
C1	Demonstrate knowledge and understanding of essential facts, concepts, principles and theories: Major aspects of chemical terminology, nomenclature, units and unit conversions.
C2	Demonstrate knowledge and understanding of essential facts, concepts, principles and theories: types of chemical reactions and its main characteristics
C5	Demonstrate knowledge and understanding of essential facts, concepts, principles and theories: Characteristics of the different states of matter and the theories used to describe them
C9	Demonstrate knowledge and understanding of essential facts, concepts, principles and theories: characteristic properties of the elements and their compounds, including group relationships and variations in the periodic table
C12	Demonstrate knowledge and understanding of essential facts, concepts, principles and theories: structural features of chemical elements and their compounds, including stereochemistry
C19	Apply knowledge and understanding to solve basic problems of quantitative and qualitative nature
D1	Communicate orally and in writing in at least one of the official languages of the University
D3	Learn independently
D4	Search and manage information from different sources
D6	Use mathematics, including error analysis, estimates of orders of magnitude, correct use of units and data representations
D7	Apply theoretical knowledge in practice
D8	Teamwork
D9	Work independently
D12	Plan and manage time properly
D13	Make decisions
D14	Analyze and synthesize information and draw conclusions
D15	Evaluate critically and constructively the environment and oneself

Learning outcomes

Expected results from this subject	Training and Learning Results
------------------------------------	-------------------------------

Interpret the functions of radial distribution and the angular representations of the s, p, d and f orbitals. Describe the configuration in the fundamental state of atoms and ions. Justify the variations of different atomic parameters along the Periodic Table. Interpret the electronegativity and the polarizability of an atom.	C5	D1
	C9	D3
	C19	D4
		D6
		D7
		D8
		D9
		D12
		D13
		D14
Recognize the atomic orbitals involved in a bonding. Build diagrams of OM for diatomic molecules and deduce properties of the bonding. Define overlap integral. Apply the method of hybridization to explain the bonding in simple molecules.	C5	D1
	C19	D3
		D4
		D7
		D8
		D9
		D12
		D13
		D14
Describe the state of aggregation of the elements and his behaviour in front of oxygen and water. Describe the natural resources of the elements and some methods of obtaining.	C5	D1
	C9	D3
		D4
		D7
		D8
		D9
		D12
		D14
Use the models of bonding to explain the structure of the main functional groups. Relate its structure with its macroscopic properties.	C1	D1
	C9	D3
		D4
		D7
		D8
		D9
		D12
		D14
Identify the acidic protons in an Brönsted acid. Classify the Brönsted acids. Predict the acidity and basicity of organic compounds. Identify acids and bases of Lewis and types of acid-base reactions. Identify acids and bases as hard or soft and explain its interaction.	C1	D1
	C2	D3
	C19	D4
		D7
		D8
		D9
		D12
		D14
Represent the three-dimensional structure of organic molecules. Apply the principles of stereochemistry. Determine the absolute configuration. Apply the nomenclatures R/S and Z/Y. Explain the bonding solids. Relate structure and properties in amorphous solids. Describe the superconductivity. Interpret one model structure. Predict the coordination number in function of the relation of ionic radii. Use the cycle of Born-Haber to determine the lattice enthalpy.	C1	
	C12	
	C5	D1
	C19	D3
		D4
		D7
		D8
		D9
		D12
		D14
Define the standard potentials of reduction. Calculate the variation of energy of Gibbs in a redox reaction. Explain an electrochemical cell. Predict the products and its quantities in a electrolysis.	C1	D1
	C19	D3
		D4
		D7
		D8
		D9
		D12
		D14

Characterize the types of radiation in a radioactive disintegration. Write nuclear reactions.
Calculate the nuclear binding energy and the half life of an isotope. Describe the reactions in nuclear chain. Enumerate examples of the use of radioisotopes.

C1
C19
D1
D3
D4
D7
D8
D9
D12
D14

Contents

Topic	
Subject 1: Structure of matter	Structure of the hydrogenic atoms. Polyelectronic atoms. Atomic parameters. Lanthanide contraction. Electronegativity. Polarizability.
Subject 2: Chemical bonding	Theory of OM. Types of orbital: sigma, pi, delta. Diagram of energies for diatomic homo- and heteronuclear molecules.
Subject 3: Organic Compounds and functional groups	Structure and geometry. Approach and nomenclature of organic compounds. Physical properties.
Subject 4: Isomery	Geometrical isomery. Conformational stereoisomery. Configurational stereoisomery.
Subject 5: Solids	Structure of the simple solids. Structure of the metals. Alloys. Metallic bonding. Semiconductors. Ionic solids. Energetic aspects.
Subject 6: Acid-Base properties of the main group elements and their compounds	Brönsted acids and bases. Lewis acids and bases.
Subject 7: Redox properties of the main group elements and their compounds	Oxidants and reductants. Nerst Equation.
Subject 8: Electrochemistry	Concentration cells. Batteries. Fuel cells. Electrolysis. Commercial electrolytic processes. Corrosion.
Subject 9: Nuclear chemistry	Nuclear reactions. Radioactive disintegration. Artificial transmutations. Nuclear fission. Nuclear fusion. Nuclear radiation. Applications of the radioactivity.

Planning

	Class hours	Hours outside the classroom	Total hours
Lecturing	26	38	64
Seminars	26	40	66
Essay questions exam	3	11	14
Essay questions exam	2	4	6

*The information in the planning table is for guidance only and does not take into account the heterogeneity of the students.

Methodologies

	Description
Lecturing	In these sessions, we present the general aspects of the program
Seminars	This teaching activity will be employed to solve some problems or exercises proposed related with the matter. These exercises will be delivered previously to the student through the platform Tem@ expecting that the student work them. In these sessions, we can collect questions or short problems to control the progress of the students.

Personalized attention

Methodologies Description

Seminars	During all the educational period the students will be able to consult all type of doubts related with the matter. These queries will attend so much in schedules of tutorials as of seminars.
----------	--

Assessment

Description	Qualification Training and Learning Results
-------------	---

Seminars	Students attitude and participation in seminar classes will be valued. Short questions and hand-made problems will be also proposed to track students' progress. Grading in this section will be only considered if students reach a mark equal or above 5/10 in the written exams.	15	C1 C2 C5 C9 C12 C19	D1 D3 D4 D6 D7 D8 D9 D12 D13 D14 D15
Essay questions exam	A final written test will be proposed to evaluate the acquired competences. To) If the first written test was passed, the final examination will be dedicated to the matter explained since then. b) Those students who have not passed the first written test will have to examine of all the matter. In this case, the weight of the final examination in the qualification will be the sum of the corresponding to both written tests.	40	C1 C2 C5 C9 C12 C19	
Essay questions exam	At half-semester, a first written test will be proposed about the matter explained so far. This exam will be eliminatory of matter in the final test if students reach a mark equal or above 5/10. Those students not reaching this mark will have to repeat this part of the examination in the final written test.	45	C1 C2 C5 C9 C12 C19	

Other comments on the Evaluation

Students must attend all tests performed along the course. The participation in the evaluation activities throughout the semester or in some of the assessment tests involve the condition of Presented and therefore the student will be graded.

Indicate that the final grade of the matter will be:

- That obtained with the continuous evaluation (15% seminars + 45% first written test + 40% final written test) for those students that reach an equal or upper punctuation to 5 points on 10 in the first written test. The assistance to the two written examinations is compulsory.
- That obtained only in the final written test after examining of all the matter for those students that have not reached an equal or upper punctuation to 5 points on 10 in the first proof written, that is, 15% seminars + 85% final written test.

Assessment in July: The same rules are applied.

Sources of information

Basic Bibliography

- Chang, R. and Goldsby, K. A., **Química**, 12ª, McGrawHill: Mexico, 2017
- Petrucchi, R.A. et al., **Química general: Principios y aplicaciones modernas.**, 11ª, Madrid: Pearson Educación, D.L., 2017
- Whitten, K.W., **Química**, 10ª, Cengage Learning, 2015
- Brown, T.L.; Lemay, H.E.; Bursten, B.E.; Murphy, C.J.; Woodward, P.M., **Química. La ciencia central.**, 12ª, Pearson: Naucalpan (Mejico), 2014
- Peterson, W. R., **Nomenclatura de las sustancias químicas.**, 4ª, Barcelona: Reverté, D.L., 2016
- Quiñoá, E. e Riguera, R., **Nomenclatura y representación de los compuestos orgánicos.**, 2ª, McGraw Hill Interamericana, 2005

Complementary Bibliography

- Frenking, G. and Shaik, S., **The Chemical bond.**, Weinheim: wiley-VCH, 2014
- Tan, J. and Chan K.S., **Understanding Advanced Physical Inorganic Chemistry.**, World Scientific Publishing, Singapore, 2017
- Pfennig, B.W., **Principles of Inorganic Chemistry.**, 1ª, Wiley, 2015
- Jr Wade, L.G., **Química Orgánica.**, 7ª, Pearson-Educación de México, 2012
- Carey, F., **Química Orgánica.**, 9ª, McGraw Hill: Interamericana, 2014
- Yurkanis, B.P., **Química Orgánica.**, 9ª, Pearson-Prentice Hall, 2008

Recommendations

Subjects that continue the syllabus

- Physical chemistry I/V11G200V01303
- Inorganic chemistry I/V11G200V01404
- Organic chemistry I/V11G200V01304

Subjects that are recommended to be taken simultaneously

Physics: Physics 2/V11G200V01201
Geology: Geology/V11G200V01205
Mathematics: Mathematics 2/V11G200V01203
Chemistry, physics and geology: Integrated laboratory 2/V11G200V01202

Subjects that it is recommended to have taken before

Biology: Biology/V11G200V01101
Physics: Physics I/V11G200V01102
Mathematics: Mathematics 1/V11G200V01104
Chemistry, physics and biology: Integrated laboratory 1/V11G200V01103
Chemistry: Chemistry 1/V11G200V01105
